

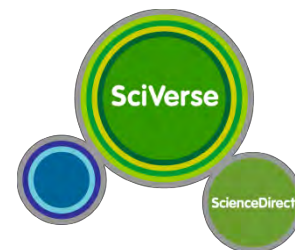
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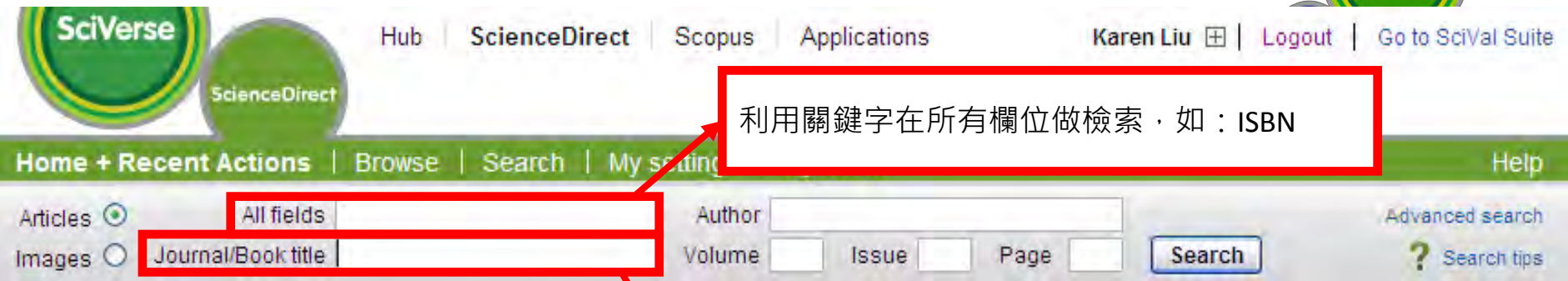
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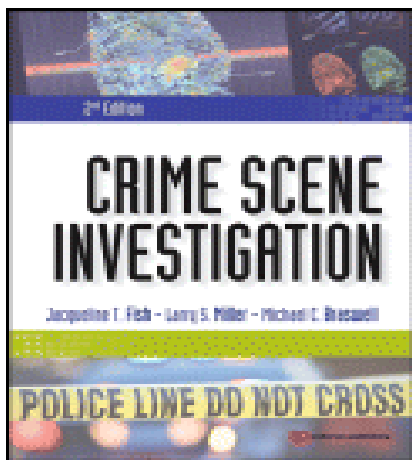


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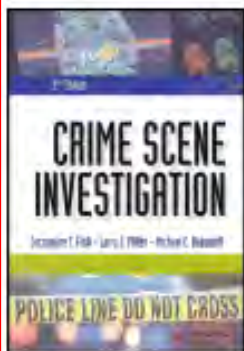
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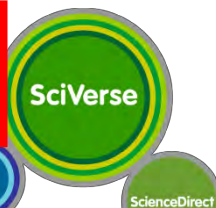
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Descriptive Inorganic Chemistry (Second Edition)

2010, Pages 63–90



CHAPTER 3 – Ionic Bonding, Crystals, and Intermolecular Forces

James E. House, Kathleen A. House

Illinois Wesleyan University, Bloomington, Illinois

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Although molecules are held together by bonds that are predominantly covalent, many substances are made up of ions that are arranged in a crystal lattice. These materials are held together in the solid state by forces that are essentially electrostatic in character. In some cases, the forces arise from the transfer of electrons between atoms to produce ionic materials. However, in most cases the ions are somewhat polarizable (especially anions), so the ions have distorted structures that represent some degree of electron sharing. As a result, many of the forces in crystals that are normally considered to be "ionic" may be appreciably less than completely ionic. This fact should be kept in mind as the principles of ionic bonding are discussed.

To predict and correlate the properties of inorganic compounds, it is essential to realize that there are other forces that are important in chemistry. These are forces that exist *between* molecules that can have a great effect on physical and chemical properties. If the molecules are polar, there will be an electrostatic attraction between them as a result of dipole-dipole forces. The hydrogen bond is a special type of this kind of interaction. Additionally, even though molecules may not be polar, there are weak forces holding the molecules together in the solid (e.g., solid CO₂) and liquid states. An understanding of these intermolecular forces is necessary to understand the chemical and physical behavior of all materials regardless of whether they are of inorganic, organic, or biological origin.

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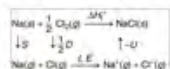


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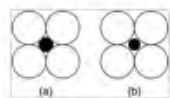
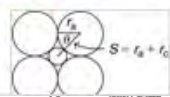


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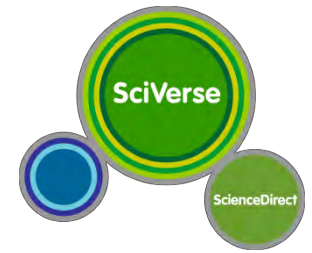
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